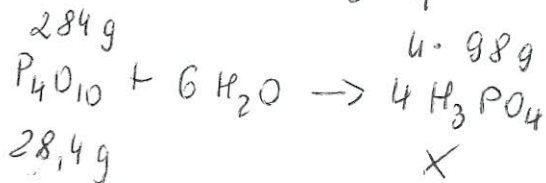


$$m_{P_4O_{10}} = 28,4g$$

$$M_{P_4O_{10}} = 4 \cdot 31 + 10 \cdot 16 = 284g/mol$$

$$m_{H_2O} = 82g$$

$$M_{H_3PO_4} = 3 + 31 + 4 \cdot 16 = 98g/mol$$



$$284g P_4O_{10} \rightarrow 4 \cdot 98g H_3PO_4$$

$$28,4g P_4O_{10} \rightarrow x$$

$$x = 4 \cdot 9,8 = 39,2 H_3PO_4$$

$$C_p = \frac{\text{masa kwasu}}{\text{masa catkowita}} \cdot 100\% = \frac{39,2g}{110,4g} \cdot 100\% = 35,5\%$$

Stężenie % kwasu - 35,5%